Name \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Naming covalent compounds

Date \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Period \_\_\_\_\_\_\_\_\_\_\_\_\_\_

**Directions:** Name the following covalent compounds using the following rules

**Rules for naming covalent compounds**

1. The first element in the formula is named first, and the full element name is used
2. The second element is named as though it were an anion *(-ide ending)*
3. Prefixes are used to denote the numbers of atoms present.
4. The prefix mono is *never* used for naming the first element

**Table of Prefixe**

|  |  |  |  |
| --- | --- | --- | --- |
| Number indicated | Prefix | Number indicated | Prefix |
| 1 |  | 6 |  |
| 2 |  | 7 |  |
| 3 |  | 8 |  |
| 4 |  | 9 |  |
| 5 |  | 10 |  |

|  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- |
| Compound  | First element name | Prefix for first element | Second element name | Prefix for second element | Compound name |
| NO2 |  |  |  |  |  |
| PCl5 |  |  |  |  |  |
| P4O6 |  |  |  |  |  |
| SF6 |  |  |  |  |  |
| SO3 |  |  |  |  |  |
| SO2 |  |  |  |  |  |
| N2O3 |  |  |  |  |  |
| CO |  |  |  |  |  |
| CO2 |  |  |  |  |  |
| SiO2 |  |  |  |  |  |
| O2F2 |  |  |  |  |  |
| XeF6 |  |  |  |  |  |
| CBr4 |  |  |  |  |  |
|  |  |  |  |  |  |

**Directions: Write for formulas for the following compounds**

|  |  |
| --- | --- |
| **Compound** | **Chemical formula** |
| iodine monochloride |  |
| nitrogen trichloride |  |
| silicon dioxide |  |
| germanium tetrahydride |  |
| dinitrogen tetrabromide |  |
| diphosphorous pentasulfide |  |
| Selenium dioxide |  |
| Silicon dioxide |  |
| dihydrogen oxide |  |
| nitrogen trihydride |  |